



Ain Shams University Faculty of Education Physics Department

Theoretical Investigation of Hydrogen Storage Capacity in Defected Carbon Materials

THESIS

Submitted in partial fulfillment of the requirement for the Master degree of Teacher Preparation in Science (Physics)

By

Mohamed Ahmed Farea Mohsen Al Khateeb

To Physics Department Faculty of Education Ain Shams University

2011

بسم الله الرحمن الرحيم

{قالوا سبحانك لا علم لنا حيناً إنك أنه العليم المكيم}

صدق الله العظيم



Ain Shams University Faculty of Education Physics Department

Researcher Name: Mohamed Ahmed Farea Mohsen Al-Khateeb

Title of the thesis: Theoretical Investigation of Hydrogen Storage Capacity in Defected Carbon Materials

Submitted to: Physics Department, Faculty of Education, Ain Shams University

Supervisors:

- 1- Prof. Dr. Mohamed Ahmed Kamel
- 2- Dr. Khaled Mahmoud Ahmed Eid
- 3- Dr. Ahlam Abd El-Monem Ahmed El-Barbary

Approval Sheet

Title :" Theoretical Investigation of Hydrogen Storage

Capacity in Defected Carbon Materials"

Candidate: Mohamed Ahmed Farea Mohsen Al-Khateeb

Degree: Master degree of Teacher Preparation in Science

(Physics)

Board of Advisors

Approved by

1. Prof. Dr/ Mohamed Ahmed Kamel

Prof. of Theoretical Physics Faculty of Education, Ain Shams University

2. Dr/ Khaled Mahmoud Ahmed Eid

Associate Professor of Theoretical Physics Faculty of Education, Ain Shams University

3. Dr/ Ahlam Abd El-Monem Ahmed El-Barbary

Lecturer of Theoretical Physics Faculty of Education, Ain Shams University

Date of presentation: / / 2011

Post graduate studies:

Stamp: / / Date of approval: / / 2011

Approval of Faculty Council: / / 2011

Approval of University Council: / / 2011



جامعة عين شمس كلية التربيــة قسـم الفيزياء

دراسة نظرية لسعة تخزين الهيدروجين في المواد الكربونيه المشوهة

رسالة مقدمة للحصول على ماجستير إعداد المعلم في العلوم (فيزياء)

مقدمة من محمد أحمد فارع محسن الخطيب

إلى

قسم الفيزياء كلية التربية جامعة عين شمس

7.11



جامعة عين شمس كلية التربية قسم الفيزياء

رسالة ماجستير

اسم الطالب : محمد أحمد فارع محسن الخطيب

عنوان الرسالة المواد الكربونيه الميدروجين في المواد الكربونيه المشوهة

اسم الدرجة : ماجستير إعداد المعلم في العلوم (فيزياء)

لجنة الإشراف:

١ ـ أ. د. محمد أحمد كامل

أستاذ الفيزياء النظرية- كلية التربية- جامعة عين شمس

٢ - د. خالد محمود احمد عيد

أستاذ مساعد الفيزياء النظرية- كلية التربية- جامعة عين شمس

٣- د. أحلام عبد المنعم أحمد البربري

مدرس الفيزياء النظرية- كلية التربية- جامعة عين شمس

تاريخ البحث: / / ٢٠١١

الدراسات العليا ختم الإجازة

أجيزت الرسالة بتاريخ / ۲۰۱۱/ موافقة مجلس الجامعة / ۲۰۱۱/

عتم الإِجارة

موافقة مجلس الكلية / ۲۰۱۱/



عين شمس كلية التربية قسم الفيزياء

صفحة العنوان

 $oldsymbol{ol}oldsymbol{oldsymbol{oldsymbol{ol{ol}}}}}}}}}}}}}}}}}}}$

سم الطالب: محمد أحمد فارع محسن الخطيب

الدرجة العلمية: الماجست يرفى اعداد المعلم في الفيزياء

القسم التابع له: الفيزياء

اسم الكلية : التربية

الجامعة : عين شمس

سنة التخرج: ٢٠٠١



امعة عين شمس كلية التربية قسم الفيزياء

شکر و تقدیر

اشكر السادة الأساتذة الذين قاموا بالإشراف على هذة الرسالة و هم:

١- أ. د/ محمد أحمد كامل
 أستاذ الفيزياء النظرية- كلية التربية- جامعة عين شمس

ـ د/ خالد محمود احمد عيد أستاذ مساعد الفيزياءالنظرية ـ كلية التربية ـ جامعة عين شمس

٣- د/ أحلام عبد المنعم أحمد البربري
 مدرس الفيزياء النظرية- كلية التربية- جامعة عين شمس

	١.	11
c	1	χı

أهدي هذا العمل إلى:

وطنــــي الذي ليس كمثله وطن وفاءً بالجميل،
روح أمــي الطاهرة دعاءً ورحمة،
روح والــدي الطاهرة دعاءً ورحمة،
زوجتــي الغالية، واولادي حبا، وتقديراً،
كل من اسدل لي معروفاً وانار لي الطريق عرفانا ً بالجميل،
الباحثون في محراب العلم دعاءً، وتوفيقاً

Abstract

This work is classified as follows:

Hydrogen is abundant, uniformly distributed throughout the Earth's surface and its oxidation product (water) is environmentally benign. Owing to these features, it is considered as an ideal synthetic fuel for new energetic matrix (renewable, secure and environmentally friendly) that could allow a sustainable future development. However, for this prospect to become a reality, efficient ways to produce, transport and store hydrogen still need to be developed. In this thesis, theoretical studies of a number of potential hydrogen storage materials have been performed using Hartree Fock (HF) and Density Functional (DFT) theories with applying 3-21G and 6-31G basis sets. The hydrogen storage in two different types of carbon materials, pristine- graphite and carbon (5.5) nanotubes, will be investigated. Then, we will explore the role of defects and doping on the cavity of hydrogen storage. Therefore, we will study isolated, mono- and di- vacancy defected graphite and (5.5) carbon nanotubes. The effect of boron and nitrogen doping in graphite and (5.5) carbon nanotubes will be also studied. Mulliken analysis, electron distribution and molecular orbitals will be applied to analyze the obtained results.

	Page
Acknowledgement	i
List of Tables	iii
List of Figures	vi
Abbreviations	xi
AbstractSummary	xii :::
Summary	xiii
Chapter I	
<u>Introduction</u>	
1.1 Current Situation of Energy Sources	1
1.2 Hydrogen as Clean, Efficient, Renewable Energy Carrier	2
1.3 Scientific Challenges And Fundamental Research Needs Of On-board Hydrogen Storage	4
1.4 Generating Green Electricity	8
1.5 Carbon	11
1.5.1 Carbon Materials	13
Graphites	13
Physical Properties of Graphites	15
Fullerenes	16
Physical Properties of Fullerenes Nanotubes	17
Physical Properties of Nanotube	18 20
1.6 Defects	22
1.61 Point Defects	22
1.6.2 Line Defects	24
Chapter II	
Theoretical background	
2.1 The Electronic Problem	25
2.1.1TheBorn-OppenheimerApproximation	26

2.1.2 The Antisymmetric or Pauli Exclusion Principl	27
2.2 Orbital and Slater Determinants	28
2.2.1 Spin Orbitals and Spatial Orbitals	28
2.2.2 The Hartree Product.	29
2.2.3 Slater Determinants	31
2.3 The Hartree-Fock Theory	32
2.3.1 The Hartree-Fock Equations	32
2.3.2 The Fock operator	35
2.4 The Density Functional Theory	26
2.4.1 The Local Density Approximation	36 38
2.5 The <i>ab Initio</i> Modeling Program Packages	39
2.5.1 Gaussian 03 Package	39
2.5.2Basis Sets	40
<u>Chapter III</u>	
Literature review	
3.1 Literature Review on Hydrogen Storage	45
3.1.1 Hydrogen Storage in Graphite	46
3.2 Literature Review on Hydrogen Storage in	
Carbon Nanotube	48
<u>Chapter IV</u>	
The role of irradiation in graphite for hydrogen	
<u>storage</u>	
4.1 Introduction	52 53
4.2 Methods	53 54
4.3 Results and Discussion	54 54
4.3.1 Ab initio calculations on graphite	54 55
4.3.2. Chemisorption of H atom on the pristine graphite	33
for models A and B	58
4.3.3 Chemisorption of H atom on the vacancy defected	20
surface	58
4.3.3.1 IVIOHO-VACANCY GEIECT	

4.3.3.2. Di-vacancy defect	
4.3.4 Chemisorption of H atom on the nitrogen doped	
surface	
4.3.4.1 Doped graphite surface with one nitrogen atom	m
4.3.4.2 Doped graphite surface with two nitrogen	
atoms	
4.3.5. Heats of adsorption or bond energies	
4.3.6. Mulliken Analysis	
4.3.6.1 Pristine graphite	
4.3.6.2 Mono-vacancy defect	
4.3.6.3 Di-vacancy defect	
4.3.6.4 Doped graphite surface with one nitrogen atom	n.
4.3.6.5 For doped graphite surface with two nitrogen	
atoms	
4.3.7. Molecular Orbitals	
4.3.7.1 Pristine graphite	
4.3.7.2 Mono-vacancy defect	
4.3.7.4 Di-vacancy defect	
4.3.7.4 Doped graphite surface with one nitrogen aton	
4.3.7.5Doped graphite surface with two nitrogen atom	IS
4.4 Conclusion	
Chapter V	
Vacancy and Heteroatoms Defects in Carbo	n
Nanotubes for Hydrogen Storage	<u></u>
Transcribed for Hydrogen Storage	
5.1 Introduction	
5.2 Methods	
5.3 Results and Discussion	
5.3.1 Ab initio calculations of CNT without hydroge	-n
storage	/11
5.3.1.1 Ab initio calculations on pristine CNT	
5.3.1.2 Ab initio calculations on mono-vacancy defect	
5.3.1.3 Ab initio calculations on di-vacancy defect	
The state of the s	. •

5.3.1.4 Ab initio calculations on isolated mono-vacancies	95
5.3.1.5 Ab initio calculations on nitrogen-doped	96
5.3.1.6 Ab initio calculations on boron-doped	98
5.3.1.7 Ab initio calculations on doping B-N	100
5.3.1.8 Ab initio calculations on doping N_2 - and B_2 -	101
mono vacancy defect	
5.3.2 Ab initio calculations of hydrogen storage inside CNT	102
5.3.2.1 Adsorption of H ₂ molecules inside the pristine	102
CNT	102
5.3.2.3 Adsorption of H ₂ molecules inside vacancy CNT. 5.3.2.3 Adsorption of H ₂ molecules inside doped CNT	103
5.3.3 Ab initio calculations of hydrogen storage outside	
CNT	107
5.3.3.1 Adsorption of H ₂ molecules outside the pristine	
CNT	107
5.3.3.2 Adsorption of H ₂ molecules outside mono-	108
vacancy defect of (5,5) CNT	
5.3.3.3 Adsorption of H2 molecules on di-vacancy	100
defect of (5,5) CNT	109
5.3.3.4 Adsorption of H ₂ molecules outside isolated	
mono-vacancy defected of (5,5) CNT	110
5.3.3.5 Adsorption of H ₂ molecules outside N-	
doped defect of (5,5) CNT	111
5.3.3.6 Adsorption of H ₂ molecules outside B-doped	
defect of (5,5) CNT	115
5.3.3.7 Adsorption of H ₂ molecules outside BN-	119
doped defect of (5,5) CNT	11)
5.4 Mulliken analysis and Molecular orbitals	122
Conclusions	127
References	128
Arabic Summary	

List of Figures

List of Figures

	Page
Figure(1.1): (a) a fule cell generates electricity by reacting oxygen, hydrogen ions and electrons. (b) the hydrogen ions generated at the anode are swept through a membrane to the cathode. This membrane is impermeable to electrons generated at the anode.	9
Figure 1.2 Structures of (a) hexagonal form and (b)	
rhombohedral form of graphite.	14
Figure 1.3 C ₆₀ bucky- ball hollow sphere $1 \approx 10$ -9m (1nm) in	17
diameter.	
Figure 1.4 Structures of Carbon Nanotubes.	20
Figure 1.5 Point Defects.	23
Figure 1.6 Line defects (Dislocations)	24
Figure 4.1 : (a) The relaxed structure of pristine graphite C_{18} ,	55
and (b) its electron density distribution.	
Figure 4.2: Model A, the relaxed structure of hydrogen atom chemisorbed on the top of carbon atom of pristine graphite C_{18} (a) side view (b) top view and (c) its electron density distribution.	57
Figure 4.3: Model B, the relaxed structure of hydrogen atom	
chemisorbed on the center of hexagon of pristine graphite C_{18} (a) side view (b) top view and (c) its electron density	57
distribution.	
Figure 4.4 : Model C, the relaxed structure of hydrogen atom chemisorbed on monovacancy defect of graphite, C_{17} H (a) side view (b) top view and (c) its electron density distribution.	60
Figure 4.5 : Model D, the relaxed structure of hydrogen atom chemisorbed on monovacancy defect of graphite, $C_{17}H$ (a) side view (b) top view and (c) its electron density distribution	61
Figure 4.6 : Model E, the relaxed structure of hydrogen atom chemisorbed on monovacancy defect of graphite, $C_{17}H$ (a) side view (b) top view and (c) its electron density distribution.	61
Figure 4.7 : Model F, the relaxed structure of hydrogen atom chemisorbed on di-vacancy defect of graphite, C ₁₆ H (a) side view (b) top view and (c) its electron density distribution	64