Ain Shams University Faculty of Science Chemistry Department



Investigation and Sintering of Nano-manganite Ceramics Powders Prepared Via Different Methods Thesis Submitted By Rehab El-Sayed Abd-Alaziz Negida (M. Sc. in Chemistry)

For Fulfillment of the Degree of Ph. D. in Chemistry

Chemistry Department Faculty of Science Ain Shams University

Cairo, Egypt 2020

Ain Shams University Faculty of Science Chemistry Department



Investigation and Sintering of Nanomanganite Ceramics Powders Prepared Via Different Methods

Thesis Submitted by

Rehab El-Sayed Abd-Alaziz Negida

(M. Sc. in Chemistry)
For Fulfillment of the Degree of Ph. D. in Chemistry
Supervised by

Prof. Dr. Eisa El-Sayed Hekal

Professor of Chemistry, Department of Chemistry, Faculty of Science, Ain Shams University.

Prof. Dr. Lame Gendy Gerges

Professor of Chemistry and Technology of Refractories, Department of Refractories, ceramics and Building Materials, National Research Centre.

Prof. Dr. Mahmmoud Farag Mahmmoud Zawrah

Professor of Chemistry of Materials and Nanotechnology, Department of Refractories, Ceramics and Building Materials, National Research Centre.

Dr. Reham Mostafa Mohammed Khattab

Assistant Prof. of Ceramics, Department of Refractories, Ceramics and Building Materials, National Research Centre.

2020

ACKNOWLEDGEMENTS

My deep gratefulness, thankful as to the merciful "Allah" who gave me everything I have as well as the ability and patience for accomplishing this work.

I would like to express my deep appreciation and my deep gratitude to **Prof. Dr. to Eisa El-sayed Hekal**, Professor of Chemistry, Faculty of Science, Ain Shams University for sponsoring the work, kind supervision, continuous encouragement, valuable advice, critical reading of the manuscript. The guidance and support especially during the latest phase of the present work.

I would like to express my cardiac thanks and gratitude to **Prof. Dr. Mahmoud Farag Mahmoud Zawrah**, Professor of Chemistry of Materials and Nanotechnology, Department of Refractories, Ceramics and Building Materials, National Research Centre, for suggesting the point, following up all stages as well as his critical writing, reading and revising of the thesis up to its final form. I will never forget his support and continuous encouragement.

I wish to express my gratitude to **Prof. Dr. Lame Gendy Gerges**, Professor of Chemistry and Technology of Refractory, Department of Refractories, Ceramics and Building Materials, National Research Centre, for his assistance, valuable help, encouragement and advices during performing this work.

My deepest thanks and respect to **Dr. Reham Mostafa**Mohammed Khattab, Researcher assistant Prof. of Ceramics,
Department of Refractories, Ceramics and Building Materials,
National Research Centre, who taught me the experimental work,
sincere criticism, critical advice, following up all stages,
continuous encouragement and supporting during carrying out the
work in this thesis.

Last but not least, my deepest gratitude to my family, whom all I am indebted for their help and encouragement during carrying out the work.

CONTENTS

<u>Content</u> - List of tables	Page
- List of figures	
- List of abbreviations	
- Abstract	
<u>Chapter I</u>	
<u>INTRODUCTION</u>	
1. Introduction	1
1.1. Introductory Remarks	1
1.2. Structure of perovskite manganites	2
1.3. Zener double exchange (ZDE) mechanism	4
1.4. Manganese metal and minerals	6
1.5. Manganese ore deposits	8
1.6. Extraction of manganese	8
1.6.1. Bio-metallurgy / Bioleaching process	9
1.6.2. Hydrometallurgy	11
1.7. Properties of nano manganese oxide	15
1.8. Methods of preparation of perovskite Manganites	16

Content	Page
1.8.1. Hydrothermal synthesis	18
1.8.2. Mechanical activation method	19
1.8.3. Advantages of mechanical activation Techniques	21
1.9. Applications of perovskite manganites	27
1.10. Objectives and scope of the present Study <u>Chapter II</u>	31
EXPERIMENTAL	
2. Material and Experimental Methods	32
2.1. Starting materials	32
2.2. Experimental Procedures	33
2.2.1. Preparation of manganite with perovskite structure	33
2.2.1.1. Hydrothermal method	34
2.2.1.2. Mechanical activation method	36
2.2.2. Consolidation of the Synthesized Composites	40
2.2.3. Characterization of the Synthesized Powder	41
i. X-ray diffraction (XRD) analysis	41
ii. The relative density and apparent porosity	42

Content	Page
iii. scanning electron microscopy (SEM)	43
iv. Transmission electron microscopy (TEM)	43
v. Broadband dielectric spectrometer (BDS)	44
vi. VSM for magnetic properties measurements <u>Chapter III</u>	44
RESULT AND DISCUSSION	
PART 1	
3. Result And Discussion	45
3.1. Hydrothermal procedures	45
3.1.1. Phase composition of calcined perovskite powders	45
3.1.2. Morphology and particle size of calcined perovskite powders	47
3.1.3. Physical properties of sintered Perovskites	49
3.1.4. Phase composition of sintered Perovskites	52
3.1.5. Microstructure of sintered perovskites	54
3.1.6. Electrical properties of perovskites	56
3.1.7. Magnetic properties of perovskites	59

<u>Content</u>	Page
3.2. Mechanical preparation method	62
3.2.1. Composition manganese ore and extracted MnO ₂	62
3.2.2. Composition and morphology of prepared perovskite manganites	64
3.2.2.1. Phase composition of La _{0.1} Ca _{0.9} MnO ₃ and La _{0.1} Sr _{0.9} MnO ₃	64
3.2.2.2. Morphology of prepared La _{0.1} Ca _{0.9} MnO ₃ and	66
La _{0.1} Sr _{0.9} MnO ₃ powders 3.2.3. Phase identification of sintered perovskites	68
3.2.4. Physical properties of sintered perovskites	70
3.2.5. Microstructure of sintered perovskites	73
3.2.6. Electrical resistivity of sintered perovskites	76
3.2.7. Magnetic properties of sintered perovskites <u>Chapter IV</u>	79
SUMMARY AND CONCLUSIONS	83
REFERENCES ARABIC SUMMARY	89

List of Abbreviations

■ **ABO**₃ Perovskite compounds

■ LMO Lanthanum manganites

■ **LCMO** Lanthanum calcium manganites

■ **LSMO** Lanthanum strontium manganites

■ **SEM** Scanning Electron Microscopy

■ **TEM** Transmission electron microscope

■ **BDS** Broadband dielectric spectrometer

■ **XRD** X-ray Diffraction Analysis

■ **XRF** X-ray fluorescernce

■ **AP** Apparent porosity

■ **BD** Bulk density

• **SOFC** Solid oxide fuel cell

Keywords:

Perovskite structure, Manganite compounds, SOFC, Nanoparticles, Hydrothermal Synthesis, Mechanical activation method.

List of Figures

Fig. No. Fig. (1)	ABO ₃ type perovskite-structure (simple cubi system)	Page
Fig. (2)	Double exchange mechanism of perovskite manganite	5
Fig. (3)	Image of autoclave used for hydrothermal preparation method	19
Fig. (4)	Planetary ball used for mechanical milling	22
Fig. (5)	Applications of mechanical alloying technique	22
Fig. (6)	Flow chart for preparation of activated and non activated manganese products	39
Fig. (7)	Compaction die set	41
Fig. (8)	XRD patterns of calcined La _{0.1} Ca _{0.9} MnO ₃ at 900 and 1100°C	46
Fig. (9)	XRD patterns of calcined La _{0.1} Sr _{0.9} MnO ₃ at 900 and 1100°C	47
Fig. (10)	TEM images of $La_{0.1}Ca_{0.9}MnO_3$ calcined at $900^{\circ}C$ (a) and its SAED (b)	48

Fig. No.		Page
Fig. (11)	TEM images of La _{0.1} Sr _{0.9} MnO ₃ calcined at 900°C (a) and its SAED (b)	49
Fig. (12)	Bulk density and apparent porosity of La _x Ca _(1-x) MnO ₃ sintered at 1250, 1350 and 1400°C	51
Fig. (13)	Bulk density and apparent porosity of La _x Sr _(1-x) MnO ₃ sintered at 1250, 1350, 1400 and 1450°C	52
Fig. (14)	XRD patterns of La _x Ca _{1-x} MnO ₃ sintered at 1400°C	53
Fig. (15)	XRD patterns of sintered La _x Sr _{1-x} MnO ₃ sintered at 1450°C	54
Fig. (16)	SEM micrographs of sintered perovskites (a) $LaMnO_3$, (b) $La_{0.5}Ca_{0.5}MnO_3$, (c) $La_{0.3}Ca_{0.7}MnO_3$, and (d) $La_{0.1}Ca_{0.9}MnO_3$	55
Fig. (17)	SEM micrographs of sintered perovskites, (a) $LaMnO_3$, (b) $La_{0.5}Sr_{0.5}MnO_3$ (c) $La_{0.3}Sr_{0.7}MnO_3$, and (d) $La_{0.1}Sr_{0.9}MnO_3$	56
Fig. (18)	Electrical resistivity of La _x Ca _(1-x) MnO ₃ sintered at 1400°C	58
Fig. (19)	Electrical resistivity of La _x Sr _(1-x) MnO ₃ sintered at 1450°C	58

Fig. No. Fig. (20)	Variation of magnetization with magnetic field for $LaMnO_3$ and $La_xCa_{1-x}MnO_3$	Page 60
Fig. (21)	Variation of magnetization with magnetic field for LaMnO ₃ and La _x Sr _{1-x} MnO ₃	61
Fig. (22)	XRD pattern of manganese ore	63
Fig. (23)	XRD patterns of extracted MnO ₂ from manganese ore	64
Fig. (24)	XRD patterns of prepared La _{0.1} Ca _{0.9} MnO ₃ calcined at (a) 900° and (b) 1100°C	65
Fig. (25)	XRD patterns of prepared La _{0.1} Sr _{0.9} MnO ₃ calcined at (a) 900° and (b) 1100°C	66
Fig. (26)	TEM images (a) and selected area electron diffraction (b) of La _{0.1} Ca _{0.9} MnO ₃ annealed at 900°C	67
Fig. (27)	TEM image (a) and selected area electron diffraction (b) of $La_{0.1}Sr_{0.9}MnO_3$ annealed at $900^{\circ}C$	67
Fig. (28)	XRD patterns of La _x Ca _{1-x} MnO ₃ sintered at 1350°C	69
Fig. (29)	XRD patterns of $La_xSr_{1-x}MnO_3$ sintered at $1350^{\circ}C$	70

Fig. No.		Page
Fig. (30)	Apparent porosity and bulk density of La _x Ca _(1-x) MnO ₃ sintered at 1100, 1250 and 1350°C	72
Fig. (31)	Apparent porosity and bulk density of $La_xSr_{(1-x)}MnO_3$ sintered at 1250, 1350 and 1400°C	73
Fig. (32)	SEM images of (a) LaMnO ₃ , (b) La _{0.1} Ca _{0.9} MnO ₃ , (c) La _{0.3} Ca _{0.7} MnO ₃ , (d) La _{0.5} Ca _{0.5} MnO ₃ sintered at 1350°C	75
Fig. (33)	SEM images of (a) LaMnO ₃ , (b) La _{0.1} Sr _{0.9} MnO ₃ , (c) La _{0.3} Sr _{0.7} MnO ₃ , (d) La _{0.5} Sr _{0.5} MnO ₃ sintered at 1400° C	76
Fig. (34)	Electrical resistivity of La _x Ca _(1-x) MnO ₃ perovskites sintered at 1350°C	78
Fig. (35)		79
Fig. (36)	Change of magnetization with magnetic field for $LaMnO_3$ and $La_xCa_{1-x}MnO_3$	81
Fig. (37)	Change of magnetization with magnetic field for $LaMnO_3$ and $La_xSr_{1-x}MnO_3$	82

List of Tables Table No. **Page** Table (1) Known Mn-minerals with their 7 compositions Different ball mill containers, 23 Table (2) ball and their materials features **33** Table (3) Chemical composition of the starting powders Table (4) Spacification of milling system 40 Table (5) Magnetization for different 61 prepared Perovskites 63 XRF results of manganese ore **Table (6)**

Abstract

During the last decade ABO₃ manganites with perovskite structure exhibit a widely studying due to its physical, magnetic and electric properties that improve their various applications like sensors, catalysts electrode materials for solid-oxide fuel cell. This work concerns on studying: 1) the effect of various cation substitution such as Ca and Sr cations on phase and microstructural of LaMnO₃ phase and 2) the role of these cation on changing electrical, physical and magnetic properties of such material. Two different ways of preparation have been applied i.e. a hydrothermal method and mechanical this work. In activation method in hydrothermal synthesis, the crystalline powder can be prepared in one step from solutions of metal salts at a temperature of 230 °C in 48 h. On other hand, the mechanical activation method are carried out by preparation the metal oxides depending on using manganese oxide that were extracted from low-grade ores by using glucose as a reducing agent in dilute nitric acid. Then the prepration of perovskite manganite samples by milling and mixing the reactant metal oxides with the extracted MnO₂ has been carried out and then added to other various oxides to prepare manganite compounds. In all the samples the Perovskite manganites have been prepared in a general formula $La_xCa_{1-x}MnO_3$ and $La_xSr_{1-x}MnO_3$ (x = 0.1, 0.3, 0.5 and 1.00) and then sintered at different temperatures ranging from 1100°C to 1450°C. The powder characteristics of metal-doped lanthanum manganites are characterized diffraction to elucidate utilizing x-ray the transformation, scanning electron microscopy (SEM), transmission electron microscopy (TEM) and selected area electron diffraction (SAED) were used to elucidate

the size and morphology of the particles and crystalline structure of these powders. For hydrothermal procedures results revealed that the maximum temperature for LCMO is 1400°c while for LSMO is 1450°c. XRD patterns show that the pervoskite phase is detected at 20 at 33-34. By addition of Ca²⁺ and Sr²⁺ the peaks are shifted to lower 20. SEM micrographs of all grains are rod like shape. The magnetic properties of hydrothermally prepared samples are detected, where the magnetization of perovskite samples is improved by the addition of Ca2+ and Sr2 and the highest results are observed for La_{0.5}Ca_{0.5}MnO₃ and La_{0.5}Sr_{0.5}MnO₃. For mechanical activation method, MnO₂ is obtained by leaching method at 95°C and characterized by XRD. Studying the bulk density and apparent porosity show that the sintering temperature of LCMO samples is 1300 °C while for LMO and LSMO is 1350°C. TEM images of an individual $La_{0.1}Ca_{0.9}MnO_3$ nanoparticles La_{0.1}Sr_{0.9}MnO₃ nanoparticles exhibits single-crystalline particles. XRD results for sintered samples prepared by mechanosynthesis show that perovskite phases prepared after 8 h of milling with sharp peaks at 20 ranged between 32-34 that shifted to lower 2θ values by the substitution of La ³⁺ with Ca²⁺ and Sr²⁺. Scanning electron microscopy (SEM) morphology for the sintered samples appears in spherical shapes. The grain sizes are different for LCMO and LSMO samples according to x values. For magnetic properties, LSMO samples have higher magnetization than LMO and LCMO samples. La_{0.5}Sr_{0.5}MnO₃ exhibits higher magnetization than the others. The electrical resistivity for both producers shows that LCMO samples are higher than LSMO samples due to their different atomic size, while LaMnO₃ exhibits the highest resistivity.